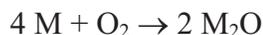


Types of Reactions

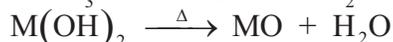
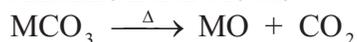
Metal Oxides

I.) Formation

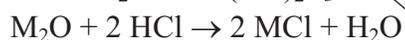
A.) Alkali Metals-very controlled amount of oxygen



B.) Alkaline Earth Metals

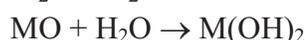


II.) Reactions

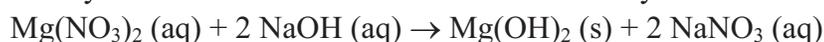


Metal Hydroxides

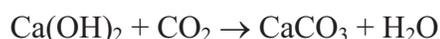
I.) Formation



Soluble hydroxide in solution with salt of insoluble hydroxide

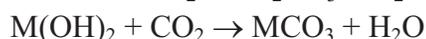
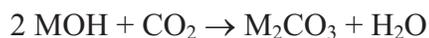
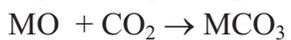
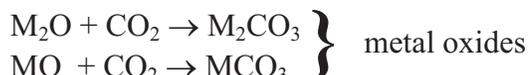


II.) Reactions



Carbonates

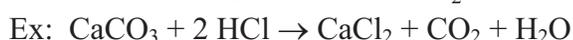
I.) Formation



II.) Reactions



Metal Carbonate + Acid \rightarrow salt + CO_2 + water



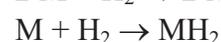
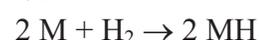
Hydrogen Carbonates

Reaction:



Hydrides

Formation

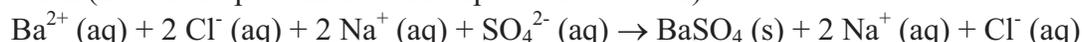


General Types of Reactions

Acid-Base Reaction: $2NaOH + H_2SO_4 \rightarrow Na_2SO_4 + 2H_2O$

Oxidation-Reduction Reaction (Single Replacement Reaction): $Mg + H_2S (g) \rightarrow MgS + H_2 (g)$

Metathesis Reaction (Double Replacement or Precipitation Reaction):



Types of Reactions

Nonmetal Oxides

Most nonmetal oxides will react with water to form oxyacids

